

Isotope Practice Worksheet

Name/period: _____

1. Sketch a block of any element on the periodic table. Label the atomic mass and atomic number. Explain what each one tells you. Use the element in your book as an example.

2. Here are three isotopes of an element: ${}_6^{12}\text{C}$ ${}_6^{13}\text{C}$ ${}_6^{14}\text{C}$
- The element is: _____
 - The number 6 refers to the _____
 - The numbers 12, 13, and 14 refer to the _____
 - How many protons and neutrons are in the first isotope? _____
 - How many protons and neutrons are in the second isotope? _____
 - How many protons and neutrons are in the third isotope? _____

3. Complete the following chart:

Isotope name	Isotope symbol	atomic #	mass #	# of protons	# of neutrons	# of electrons
92 uranium-235						
92 uranium-238						
5 boron-10						
5 boron-11						

4. Naturally occurring europium (Eu) consists of two isotopes with a mass of 151 and 153. Europium-151 has an abundance of 48.03% and Europium-153 has an abundance of 51.97%. What is the atomic mass of europium?

5. Strontium consists of four isotopes with masses of 84 (abundance 0.50%), 86 (abundance of 9.9%), 87 (abundance of 7.0%), and 88 (abundance of 82.6%). Calculate the atomic mass of strontium.